Candy Conservation of Mass

In this activity you will use skittles candy to model chemical reactions and explore the Law of conservation of mass.

1. Sort your candy into piles according to color. Whichever pile has the most pieces will represent Hydrogen (H). The next biggest pile will represent Oxygen (O), then Carbon(C), Nitrogen (N) and Sodium (Na). You will need at least 6 Hydrogen, 4 Oxygen, 2 Carbon, 2 Nitrogen, and 2 Sodium Please fill out the box below with your colors and amounts

Element	Hydrogen	Oxygen	Carbon	Nitrogen	Sodium
Color					
Number					

- 2. In this activity you will need to balance the equations to make sure that everything on the left side of the equation is equal to the right side of the equation. For example:
- $2Fe₂O₃ + 3C \longrightarrow 4Fe + 3CO₂$ Fe = 4 O = 6 C = 3 Fe = 3 C = 3
- 3. After you balance each equation on the chemical reactions sheet you will model each reaction with skittles and draw the reaction in your notebook. Start by building the

molecules on the left side of the equation, then use those same skittles to build the molecules on the right side.

4. When you have completed all of the reactions, you may eat your skittles as you answer the conclusion questions in your notebook.

Drawing:

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Reaction 2:	$\underline{}$ $H_2 + N_2 \rightarrow$	NH ₃
	H=	H=
	N=	N=

Drawing:

Reaction 3:
$$CH_4 + __O_2 \rightarrow CO_2 + __H_2O$$

 $C= C= H= H= O= O=$

Drawing:

Reaction 4: ___Na + ___H
$$_2$$
O \rightarrow ____ NaOH + H $_2$ Na= Na= H= H= O= O=

Drawing:

Reaction 5:
$$C_2H_5OH+ __O_2 \rightarrow __CO_2 + __H_2O$$

 $C= C= H= H= O= O= O= CO_2 + CO_2$

Drawing: